A Resource utilization method for volatile organic compounds emission from the semiconductor industry: Selective catalytic oxidation of isopropanol to acetone Over Au/α-Fe₂O₃ nanosheets

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ABSTRACT

We prepare α-Fe₂O₃ nanosheet supported 0.38, 0.81, and 1.36 wt% Au (average particle size = 4.0 nm) nanocatalysts, and investigate their performance and mechanism for the selective catalytic oxidation of isopropanol to acetone. In the presence of 1.2 vol% isopropanol and 40 vol% O₂, 1.36 wt% Au/α-Fe₂O₃ exhibits excellent catalytic performance, due to its moderate acidic sites and better redox properties, with acetone selectivity and yield being as high as 99% and 95% at 220 °C, respectively. In addition to acetone, little propylene, acetic acid, acetaldehyde, methyl vinyl ketone, 2-butanone, isopropyl ether, isopropyl acetate, 3-penten-2-one, isopropyl acrylate, isopropyl propionate, and 2, 4-dimethylfuran are detected. The possible reaction mechanism is proposed for the selective catalytic oxidation of isopropanol to acetone over the present catalysts. We believe that the present selective catalytic oxidation method, rather than the traditional complete catalytic oxidation method, provides an alternative and economic method for VOCs emissions control.

1. Introduction

The semiconductor industry plays an important role in promoting social and economic development. However, some environmental problems caused by the large scale production of electronic products could not be overlooked. Semiconductor devices are manufactured via the complex production processes, involving the high frequency generation and emission of hazardous pollutants. Among the volatile organic compounds (VOCs) emitted from the semiconductor industry, in addition to little benzene, toluene, xylene, ethyl acetate, and styrene, isopropanol and acetone are the major components [1,2], because isopropanol and acetone are widely used as cleaning agents. Under sunlight irradiation, ozone and secondary organic aerosols might be produced by the photochemical reaction between VOCs and nitrogen oxides, resulting in elevated levels of atmospheric oxidizing capacity and fine particulate matter concentration. Human body to prolonged exposure will raise the carcinogenic risk. Therefore, the purification of VOCs has attracted more and more attention [3–5].

Catalytic oxidation is a traditional method for VOCs purification, and VOCs could be completely oxidized into CO₂ and H₂O without secondary pollution [6,7]. Metal oxide supported noble metal (e.g. Au, Pt and Pd [8–11]) catalysts have been extensively studied due to their excellent low temperature activities. However, control VOCs emission via the catalytic oxidation needs great investment. In order to improve the enthusiasm of the enterprise owners for environmental protection, it is highly desired to develop a novel efficient method for the treatment of air pollution with extra economic benefits.

Haruta et al. [12] found that Au nanoparticles loaded on transition metal oxides, such as TiO₂, α-Fe₂O₃, Co₃O₄, and NiO, exhibited excellent catalytic oxidation activities. Actually, a number of studies show...
that the supported Au catalysts exhibit remarkable performance in various important reactions including CO oxidation and water-gas shift reactions [13,14]. α-Fe2O3 with various morphologies such as nanodisk [15], hollow [16], and columnar superstructure [17], modified by Au nanoparticles, have been widely used as heterogeneous catalysts and in situ gas sensors. For the Au/Fe2O3 catalysts, their catalytic performance is associated with the morphologies [16] and crystal phases [18] of Fe2O3 as well as the nature of the reactants [19]. In the literature, it is rare to across the complete or selective catalytic oxidation of isopropanol over Au/α-Fe2O3.

As we know, acetone has been widely used in various fields. Up to now, there are three main methods to prepare acetone. The cost of the generation of acetone via the grain fermentation is too high. The main method for acetone preparation is the oxidation of cumene synthesized from benzene and propylene, and co-production of acetone and phenol. Selective catalytic oxidation of isopropanol to acetone is an important supplementary method. In recent years, selective catalytic oxidation of isopropanol has been extensively studied. The dehydration and hydrogenation of isopropanol produce propylene and acetone, respectively. For example, the g-C3N4/Ag/Ag3PO4 composite [20] had excellent photocatalytic activity for the selective oxidation of gaseous isopropanol, and 63 % isopropanol was mainly oxidized to acetone after 4 h irradiation. Friedel et al. [21] reported that for the selective catalytic oxidation of isopropanol over novel ex-LDH CoFe2O4, the corresponding yield of acetone and propylene was 40 % and 52 % at 90 % conversion of isopropanol. Due to the high demand of economic method for VOCs control and resource utilization, we herein prepare α-Fe2O3 nanosheet supported Au nanocatalysts, and investigate their performance and mechanism for the selective catalytic oxidation of isopropanol to acetone.

2. Experimental

2.1. Catalyst preparation

α-Fe2O3 nanosheets (NSs) are synthesized via a solvothermal method [22]. The preparation procedures are as follows: 1.638 g of FeCl3·6H2O powder is dissolved in a mixed solution of 60 mL CH3CH2OH and 4.2 mL H2O. After magnetic stirring for 0.5 h, 4.8 g of FeCl3·6H2O and 0.05 g of CH2COONa is added to the mixture. The above precursor solution is stirred for 1 h, and then transferred to autoclave and kept in an oven at 180 °C for 12 h. α-Fe2O3 NSs are obtained after centrifugal separation for five times at 8000 rpm with a mixture of ethanol and deionized water. Commercial α-Fe2O3 (analytical purity, > 99.0 %) is used as a reference catalyst.

x wt% Au/α-Fe2O3 NSs (x = 0.38, 0.81, and 1.36) is prepared by colloidal deposition method that is assisted by polyvinyl alcohol (PVA) and NaBH4 [23]. The mass ratio of Au to PVA is 1.5: 1, and the molar ratio of Au to NaBH4 is 1: 5. In an ice-water bath, 1.5 mmol/L HAuCl4 aqueous solution is added with appropriate PVA solution under stirring and dark conditions. The NaBH4 aqueous solution is quickly injected to obtain the Au nanoparticles (NPs) colloidal solution. After stirring for 0.5 h, the appropriate amount of α-Fe2O3 NSs is added to the above solution. Gold colloids could be well deposited on the surface of the support after continuous magnetic stirring for 12 h. The final catalysts are obtained after centrifugal separation, washing, drying, and calcination in a muffle furnace at 300 °C for 2 h. All the chemical reagents (analytical purity) are purchased from Sinopharm chemical reagent co. LTD and used without further purification.

2.2. Catalyst characterization

All the samples are characterized by various techniques, including X-ray diffraction (XRD), 57Fe-Mössbauer spectra, scanning electron microscopy (SEM), transmission electron microscopy (TEM), and X-ray photoelectron spectroscopy (XPS), hydrogen temperature programmed reduction (H2-TPR), thermal analyzer-gas chromatography/mass spectrometry (GC/MS), isopropanol-in situ diffuse reflectance Fourier transform infrared spectroscopy (isopropanol-in situ DRIFTS), pyridine-in situ DRIFTS, isopropanol temperature programmed desorption (isopropanol-TPD), and isopropanol temperature programmed surface reaction (isopropanol-TPSR). The detailed characterization conditions are included in the Supporting Information.

2.3. Catalytic performance evaluation

The catalytic performance evaluation for the catalysts is carried out in a tubular quartz fixed-bed reactor (i. d. = 6 mm) under 0, 25, 50, 75, or 100 % of relative humidity (RH). The reactants and products are detected by gas chromatography (GC-2010, Shimadzu) with FID detector and Stabilwax@-DB column capillary column. The catalyst (about 0.050 g, 40–60 mesh) is mixed with quartz sands at a mass ratio of 1: 5 to avoid local temperature runaway, and pretreated in 20 mL/min of O2 flow at 300 °C for 1 h. By controlling the temperature, the saturated vapor pressure of isopropanol is calculated according to the Antoine Equation, and different concentrations (0.1 vol%, 0.75 vol%, and 1.2 vol%) of isopropanol is carried out by nitrogen gas. The oxygen gas concentration is 40 vol% or 10 vol%, and the nitrogen gas is the equilibrium gas. According to the amount of catalyst, and the total flow rate of the mixture gas, the space velocity (SV) is calculated to be 20,000 mL/(g h). All the catalytic performance is performed under steady-state reaction conditions, and the data are collected for three times, with the relative standard deviation less than 0.4 %, after keeping at a given temperature for 30 min.

We also determine the relative molar correction factors according to the following procedures. Trace amounts of acetone (99.9 %) is firstly dissolved in isopropanol (99.9 %). After sampling, the corresponding peak area is detected by gas chromatography. The absolute mass correction factors are calculated according to the ratio of their mass to the peak area. On this basis, the relative mass correction factors of acetone to isopropanol could be calculated. Parallel measurements are carried out three times, and the average value is adopted. In the present study, the relative molar correction factor is 1.13. The selectivity and yield of acetone are calculated by using the relative molar correction factor (fA) of acetone to isopropanol. The calculation formulas are as follows (a–c), where A2 and A3 are the peak area of isopropanol at the inlet and outlet, and A3 is the peak area of acetone at the outlet of the reactor tube detected by GC.

(a) Isopropanol conversion = (A1 − A3)/A1 × 100 %
(b) Acetone selectivity = A3 × fA/(A1 − A2) × 100 %
(c) Acetone yield = A3 × fA/A3 × 100 %

3. Results and discussion

3.1. Crystal structure, surface area, morphology, and size distribution of commercial α-Fe2O3, α-Fe2O3 NSs, and x wt% Au/α-Fe2O3 NSs

The actual Au contents of present x wt% Au/α-Fe2O3 NSs are 0.38, 0.81, and 1.36 wt%, determined by ICP-AES (Table S1). According to the standard card (JCPDS PDF# 33−0664), we find that the XRD patterns of commercial α-Fe2O3, and the as-prepared α-Fe2O3 NSs and x wt% Au/α-Fe2O3 NSs exhibit typical characteristic of rhombohedral α-Fe2O3 (Fig. S1). Compared with commercial α-Fe2O3, α-Fe2O3 NSs and x wt% Au/α-Fe2O3 NSs exhibit wider diffraction peaks and lower relative strength, due to the increase of lattice strain and the decrease of grain size [24]. Lattice strain might cause the increase of surface defects, which benefits the reactant adsorption. The grain sizes of the α-Fe2O3 NSs supports are 35.7–43.2 nm (Table S1) for the as-obtained

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catalysts, much smaller than that for commercial α-Fe2O3. Accordingly, the surface areas of α-Fe2O3 NSs and x wt% Au/α-Fe2O3 NSs (12.6–18.4 m²/g) are much larger than that of commercial α-Fe2O3 (4.3 m²/g). The ⁵⁷Fe-Mössbauer spectra (Fig. 1) show that the peak shapes of the α-catalysts, much smaller than that for commercial α-Fe2O3. Accordingly, curve) of the spectra.

Fig. 1.

Table 1

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>Surface element composition (mol/mol)</th>
<th>Actual H₂ consumption (mmol/g)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Fe³⁺/Fe²⁺</td>
<td>Oads/Olatt</td>
</tr>
<tr>
<td>α-Fe₂O₃ NSs</td>
<td>2.67</td>
<td>0.24</td>
</tr>
<tr>
<td>0.38 wt% Au/α-Fe₂O₃ NSs</td>
<td>2.49</td>
<td>0.26</td>
</tr>
<tr>
<td>0.81 wt% Au/α-Fe₂O₃ NSs</td>
<td>2.07</td>
<td>0.34</td>
</tr>
<tr>
<td>1.36 wt% Au/α-Fe₂O₃ NSs</td>
<td>1.96</td>
<td>0.38</td>
</tr>
</tbody>
</table>

a Determined from XPS.

b Determined from H₂-TPR.

Shown in Fig. S4 are the Fe 2p, O 1s, and Au 4f XPS spectra of α-Fe₂O₃ NSs and Au/α-Fe₂O₃ NSs. In the Fe 2p XPS spectra (Fig. S4A), there are two asymmetric peaks, due to the 2p₁/₂ and 2p₃/₂ orbits, respectively. Each orbit could be divided into four parts at different binding energies (BEs). For the Fe 2p₁/₂ orbit, the signals at BEs = 709.6 and 714.3 eV, and BEs = 711.2 and 718.5 eV, could be due to Fe²⁺ and Fe³⁺, respectively. For the Fe 2p₃/₂ orbit, the signals at BEs = 722.7 and 727.1 eV, and BEs = 724.4 and 732.3 eV, could be due to Fe²⁺ and Fe³⁺ [25,26], respectively. The presence of Fe²⁺ on the surface might be partially associated with the weak reduction of Fe³⁺ by ethanol used in the preparation process. The O 1s XPS spectra (Fig. S4B) could be divided into three parts at BEs = 529.4, 531.1, and 532.9 eV, due to the surface lattice oxygen (Oₐₙₙ), surface adsorbed oxygen (Oads), and surface hydroxyl species [27], respectively. The Au 4f XPS spectra (Fig. S4C) could be divided into four parts at BEs = 83.9 and 87.6 eV, and BEs = 84.9 and 88.6 eV. The former two could be due to Au⁺ species, and the latter two could be due to Au²⁺ (e.g., Au⁰ or Au³⁺) species [15]. As shown in Table 1, the surface Fe³⁺/Fe²⁺ molar ratio decreases from 2.67 to 1.96, while the surface Au²⁺/Au⁰ molar ratio increases from 0.23 to 0.31, due to the interaction (Au⁺ + Fe³⁺ → Au²⁺ + Fe²⁺) between Au NPs and α-Fe₂O₃ NSs, with a rise in the Au loading from 0 to 1.36 wt%. Furthermore, the surface Oads/Olatt molar ratio also increases from 0.24 to 0.38.

Fig. 2 gives the H₂-TPR profiles of commercial α-Fe₂O₃, α-Fe₂O₃ NSs and Au/α-Fe₂O₃ NSs. Clearly, there are two main reduction peaks in the range of 250 – 450 °C and 450 – 750 °C over the present catalysts. For the commercial α-Fe₂O₃ and α-Fe₂O₃ NSs, the peak I corresponds to the reduction of Fe₂O₃ to Fe₃O₄, and the peak II corresponds to the reduction of Fe₂O₃ → FeO → Fe⁰ [18]. Compared with that over commercial α-Fe₂O₃, the temperature of peak I over α-Fe₂O₃ NSs decreases from 394 to 372 °C, indicating a slight improvement in low temperature reducibility. Moreover, the loading of Au NPs onto α-Fe₂O₃ NSs could further enhance the low temperature reducibility. The temperature of peak I over 0.38, 0.81, and 1.36 wt% Au/α-Fe₂O₃ NSs is 336, 299, and 267 °C, respectively. The decrease in the initial reduction temperature is associated with the interaction (Au⁺ + Fe³⁺ → Au²⁺ + Fe²⁺) between Au NPs and α-Fe₂O₃ NSs. Such interaction not only gives rise to the generation of Au²⁺, but also lengthens the distance between Fe and O atoms, and weakens the Fe–O bond. In addition to the reduction of Fe₂O₃ to Fe₃O₄, the peak I over Au/α-Fe₂O₃ NSs also corresponds to the reduction of Au³⁺ to Au⁺. The total H₂ consumption of α-Fe₂O₃ NSs and Au/α-Fe₂O₃ NSs is in the range of 18.54 – 22.12 mmol/g (Table 1). It should be noted that due to the generation of more Fe²⁺ species and surface defects, the total H₂ consumption of x wt% Au/α-Fe₂O₃ NSs is lower than that of α-Fe₂O₃ NSs, and slightly decreases with an increase in the Au loading.

In order to determine the surface acidity of commercial α-Fe₂O₃, α-
The acid amount of different catalysts determined by NH3-TPD, the amount of Lewis acid at 1450 cm\(^{-1}\) and Brønsted acid at 1540 cm\(^{-1}\) determined by pyridine adsorption.

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>T (°C)</th>
<th>Total</th>
<th>Brønsted acidic site</th>
<th>Lewis acidic site</th>
<th>Brønsted/Lewis acidic site</th>
</tr>
</thead>
<tbody>
<tr>
<td>commercial</td>
<td>427, 16</td>
<td>200</td>
<td>1.14</td>
<td>30.55</td>
<td>0.037</td>
</tr>
<tr>
<td>α-Fe(_2)O(_3)</td>
<td>517</td>
<td>350</td>
<td>0.66</td>
<td>11.45</td>
<td>0.058</td>
</tr>
<tr>
<td>α-Fe(_2)O(_3) NSs</td>
<td>300, 62</td>
<td>200</td>
<td>2.21</td>
<td>43.01</td>
<td>0.051</td>
</tr>
<tr>
<td>365</td>
<td>350</td>
<td>0.94</td>
<td>14.55</td>
<td>0.065</td>
<td></td>
</tr>
<tr>
<td>1.36 wt% Au/α-Fe(_2)O(_3) NSs</td>
<td>243, 74</td>
<td>200</td>
<td>1.98</td>
<td>105.68</td>
<td>0.019</td>
</tr>
<tr>
<td>358</td>
<td>350</td>
<td>0.89</td>
<td>37.19</td>
<td>0.024</td>
<td></td>
</tr>
</tbody>
</table>

weak and strong acidity is mainly caused by the Lewis acid [29] and Brønsted acid [30] sites, respectively. For the purpose of further determination of acidity (e.g., Brønsted acid or Lewis acid) and acid strength, we carry out the pyridine-\(\text{in situ}\) DRIFTS experiment at 200 and 350 °C (Fig. 4). Generally, the band at 1450 cm\(^{-1}\) is the characteristic absorption peak of pyridine ring on the Lewis acidic site, the band at 1490 cm\(^{-1}\) is the absorption of protonated pyridine on the Brønsted acidic site, and the band at 1601 cm\(^{-1}\) is the absorption of pyridine on the Lewis acidic site. The band at 1540 cm\(^{-1}\) is the characteristic absorption peak of pyridine on the Brønsted acidic site, the band at 1601 cm\(^{-1}\) is the absorption of pyridine on the Lewis acidic site, and the band at 1650 cm\(^{-1}\) is the absorption of protonated pyridine on the Brønsted acidic site [31]. The commercial α-Fe\(_2\)O\(_3\), α-Fe\(_2\)O\(_3\) NSs and 1.36 wt% Au/α-Fe\(_2\)O\(_3\) NSs exhibit the larger amount of Lewis acid, because Fe\(^{3+}\) could provide Lewis acidic sites. Compared with that at 200 °C, the peak intensity due to the Lewis acidic site is lower at 350 °C, indicating the Lewis acid is weak and easy to remove. Table 2 shows the acid amount of the commercial α-Fe\(_2\)O\(_3\), α-Fe\(_2\)O\(_3\) NSs and 1.36 wt% Au/α-Fe\(_2\)O\(_3\) NSs calculated on the basis of the pyridine-\(\text{in situ}\) DRIFTS experiment. It can be seen that the loading of Au onto the support greatly increases the Lewis acidic sites, and slightly decreases the Brønsted acidic sites, in consistence with the NH3-TPD results.

3.3. Performance for selective catalytic oxidation of isopropanol over commercial α-Fe\(_2\)O\(_3\), α-Fe\(_2\)O\(_3\) NSs, and x wt% Au/α-Fe\(_2\)O\(_3\) NSs

As mentioned above, isopropanol and acetone are the major components, among the VOCs emissions from the semiconductor industry. We herein investigate the catalytic performance of α-Fe\(_2\)O\(_3\) NSs and Au/α-Fe\(_2\)O\(_3\) NSs for the selective catalytic oxidation of isopropanol to acetone. Under the reactants composed of 0.1 vol% isopropanol and 40 vol% O\(_2\), with N\(_2\) as equilibrium gas, the isopropanol conversion, acetone selectivity and yield over the present catalysts as a function of the reaction temperature is shown in Fig. S5. Over α-Fe\(_2\)O\(_3\) NSs, isopropanol only begins to react with O\(_2\) at a temperature higher than 120 °C, and isopropanol conversion increases with a rise in temperature. Acetone selectivity first increases from 74 % to 93 % with a rise in temperature from 120 to 150 °C, and then slightly decreases from 93 % to 81 % with a further rise in temperature from 150 to 240 °C. Acetone yield increases with a rise in temperature below 240 °C, while greatly decreases above 240 °C, due to the significant decrease in acetone selectivity. The loading of Au NPs onto the α-Fe\(_2\)O\(_3\) NSs causes the selective catalytic oxidation of isopropanol to take place at room temperature (Fig. S5A). The trend of isopropanol conversion, acetone selectivity and yield over Au/α-Fe\(_2\)O\(_3\) NSs is similar to that over α-Fe\(_2\)O\(_3\) NSs. In addition to the low reaction temperature, another big difference is the acetone selectivity over Au/α-Fe\(_2\)O\(_3\) NSs could be kept at a high level in the range of room temperature to 170 °C. The 1.36 wt % Au/α-Fe\(_2\)O\(_3\) NSs catalyst exhibits relatively better catalytic
performance for the present reaction. At room temperature, isopropanol conversion and acetone selectivity is 36.7 % and 93.3 %, respectively. In the range of 120 − 170 °C, acetone selectivity and yield are much more than 90 %. At 150 °C, the acetone yield is up to 96.2 %, where isopropanol conversion is 96.7 %, and acetone selectivity is about 99.5 % (Table S3). The Au dispersion of 0.38 wt%, 0.81 wt%, and 1.36 wt% Au/α-Fe₂O₃ NSs, determined by the CO pulse chemisorption method, is 14.2, 19.0 and 28.6 %, respectively. Then we calculate TOF_Au, isopropanol consumption rate, and acetone formation rate at 30 °C. Although TOF_Au (17.8 × 10⁻³ s⁻¹) of the 0.38 wt% Au/α-Fe₂O₃ NSs catalyst is larger than these (9.0 × 10⁻³ s⁻¹ and 4.2 × 10⁻³ s⁻¹) of the 0.81 wt% and 1.36 wt% Au/α-Fe₂O₃ NSs catalysts, isopropanol consumption rate (4.9–8.3 × 10⁻⁸ mol g_cat⁻¹ s⁻¹) and acetone formation rate (3.8–7.8 × 10⁻⁸ mol g_cat⁻¹ s⁻¹) increase with a rise in the Au loading. It should be noted that acetone selectivity could not reach 100 %, due to the presence of little intermediate products, which will be discussed later. Under an oxygen-rich reaction conditions, the oxidative dehydrogenation of isopropanol, rather than the deep oxidation.
of isopropanol and/or acetone, mainly occurs over 1.36 wt% Au/α-
Fe2O3 NSs at a low temperature.

Theoretically, 0.1 vol% isopropanol would be converted to 0.1 vol% acetone. Hence, we also investigate the catalytic activity of 1.36 wt% Au/α-Fe2O3 NSs for the complete oxidation of acetone (Fig. S6), and find that the initial temperature required for acetone oxidation is 180 °C under the reactants consisted of 0.1 vol% acetone and 40 vol% O2, with N2 as equilibrium gas. Due to the low catalytic activity of 1.36 wt% Au/α-Fe2O3 NSs for the complete oxidation of acetone, the high selectivity and yield of acetone in the range of 120 – 170 °C are understandable, for the selective catalytic oxidation of isopropanol. Nevertheless, when the reaction temperature is high enough, acetone will undergo deep oxidation (Table S4), which is the possible reason for the acetone selectivity and yield greatly decreases above 210 °C (Fig. S5).

For comparison, we also investigate the selective catalytic oxidation of isopropanol to acetone over commercial α-Fe2O3 (Fig. S7). Unlike that over α-Fe2O3 NSs, propylene is the main product, in addition to little acetone. Propylene selectivity increases with a rise in temperature below 150 °C, and then slightly decreases with a further rise in temperature from 150 to 250 °C. The highest propylene yield is 72 %, achieved at 180 °C.

### 3.4. Possible Intermediate Products Generated from Selective Catalytic Oxidation of Isopropanol Over Commercial α-Fe2O3, α-Fe2O3 NSs, and 1.36 wt% Au/α-Fe2O3 NSs

For the purpose of understanding the selective catalytic oxidation process, we try to detect the possible intermediate products by GC–MS at various temperatures over α-Fe2O3 NSs, commercial α-Fe2O3, and 1.36 wt% Au/α-Fe2O3 NSs. During the evaluation of catalytic performance, the peaks detected by GC-2010 due to isopropanol and acetone are well separated, while these detected by GC-MS are not very well separated (Fig. 5). In addition, the merged peaks also could be observed, due to the presence of several intermediate products with similar formula weight. Under the reactants consisted of 0.1 vol% isopropanol ((CH3)2CHOH) and 40 vol% O2, in addition to the major product acetone (CH3COCH3), there are isopropyl formate (HCOOC(CH3)2), methyl vinyl ketone (CH3CHO + CH2), 2-butane (CH3CH2CH3), isopropyl ether ((CH3)2CHOCH(CH3)2), and isopropyl acetate (CH3COOCH(CH3)2) over α-Fe2O3 NSs at 710 %, 65 %, and/or 790 % (Fig. 5a). With a rise in the isopropanol concentration from 0.1 to 1.2 vol%, acetone is still the major product. However, some new signals due to acetic acid (CH3COOH), acetaldehyde (CH3CHO), and 3-penten-2-one (CH3COCH2CH3) are detected, and the intensity of the peak due to methyl vinyl ketone (CH3CHO + CH2) becomes much stronger at 790 %. We deduce that 3-penten-2-one is generated from the aldol condensation between CH3CHOH and CH3CHO. Propylene (CH2=CHCH3) is not detected over α-Fe2O3 NSs, possibly due to low content and weak adsorption in the enrichment tube (Fig. 5b). As mentioned above, unlike that over α-Fe2O3 NSs, propylene (CH2=CHCH3) is the main product, in addition to little acetone over commercial α-Fe2O3 (Fig. S7). From Fig. 5c, in addition to acetone and isopropyl ether ((CH3)2CHOHCH(CH3)2), significant signal due to propylene is clearly observed at 760 % over commercial α-Fe2O3. In the presence of 1.2 vol% isopropanol (Fig. 5d), acetone is the major product, and propylene is not detected at 790 % over 1.36 wt% Au/α-Fe2O3 NSs. Methyl vinyl ketone (CH3COCH2) = CH2, 2-butenone (CH3COCH2CH3), acetic acid (CH3COOH), and isopropyl acetate (CH3COOCH(CH3)2), isopropyl acrylate (CH2=CHCOOCH(CH3)2), isopropyl propionate (CH3CH2COOCH(CH3)2), and 2,4-dimethylfuran are also detected due to the various side reactions.

We also adopt the in situ DRIFTS technique to further investigate the selective catalytic oxidation process (Table 3 and Fig. 6). Under the reactants composed of 0.1 vol% isopropanol and 40 vol% O2, with N2 as equilibrium gas, isopropanol is adsorbed by commercial α-Fe2O3 at room temperature for 30 min, and then purged with N2 (Fig. 6a). The absorption band at 3663 cm⁻¹ is attributed to the O–H stretching vibration (ν(O–H)) of isopropanol. The asymmetric and symmetric stretching vibrations (νas andνs) of CH3 from CH3COOH are detected at 2975 and 2880 cm⁻¹. The band at 1465 cm⁻¹ and 1389 cm⁻¹ is due to the asymmetric and symmetric bending vibration (δas and δs) of CH2, respectively. The peak at 1150 cm⁻¹ is due to the skeletal C–C stretch (ν(C=C)) and at 957 cm⁻¹ is due to the CH3 rocking vibration (ρ(CH3)).

### Table 3

<table>
<thead>
<tr>
<th>Frequency (cm⁻¹)</th>
<th>Assignment</th>
<th>Group</th>
<th>Molecular Formula</th>
<th>Reference</th>
</tr>
</thead>
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<tr>
<td>3663</td>
<td>ν(O–H)</td>
<td>isopropanol</td>
<td>(CH3)2CHOH</td>
<td>[32]</td>
</tr>
<tr>
<td>2975–2963</td>
<td>νas(CH3)</td>
<td>isopropanol</td>
<td>(CH3)2CHOH</td>
<td>[40]</td>
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<tr>
<td>2880–2878</td>
<td>νs(CH3)</td>
<td>isopropanol</td>
<td>(CH3)2CHOH</td>
<td>[40]</td>
</tr>
<tr>
<td>1760–1728</td>
<td>ν(C = O)</td>
<td>aldehyde</td>
<td>carboxylic acid ester</td>
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<tr>
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<td>[6,36]</td>
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<tr>
<td>1231</td>
<td>ν(C–O–C)</td>
<td>acetate</td>
<td>CH2–COOR'</td>
<td>[38]</td>
</tr>
<tr>
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<td>propionate</td>
<td>CH2(CH2)COOH</td>
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<tr>
<td>1164</td>
<td>ν(C–C)</td>
<td>isopropyl</td>
<td>(CH3)2CHOH</td>
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<td>[40]</td>
</tr>
<tr>
<td>1129</td>
<td>ρ(CH2)</td>
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<td>(CH3)2CHOH</td>
<td>[33]</td>
</tr>
<tr>
<td>1094</td>
<td>ν(O–C–C)</td>
<td>secondary</td>
<td>CH2(CH2)COOH(CH3)2</td>
<td>[38]</td>
</tr>
<tr>
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<td>ν(C–O)</td>
<td>isopropyl</td>
<td>(CH3)2CHOH</td>
<td>[37]</td>
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<td>1015</td>
<td>ν(C–O)</td>
<td>ether</td>
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<td>R–O–R’</td>
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<tr>
<td>957</td>
<td>γ(O–H)</td>
<td>isopropanol</td>
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<td>[32]</td>
</tr>
<tr>
<td>907</td>
<td>δ(CH=CH3)</td>
<td>propylene</td>
<td>CH2=CHCH3</td>
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</table>

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intensity of the band at 1760‒1728 cm\(^{-1}\) due to the presence of aldehyde (R−CHO), the carboxylic acid (R−COOH), and ester carbonyl groups (R−COOR') [36,37,39] become strong (Fig. 6d and Fig. 6f). It further confirms the generation of 2-butanone (CH\(_3\)COCH\(_2\)CH\(_3\)), acetaldehyde (CH\(_3\)CHO), acetic acid (CH\(_3\)COOH), isopropyl acetate (CH\(_3\)COOCH\((CH_3)\)\(_2\)), and 3-penten-2-one (CH\(_3\)COCH = CHCH\(_3\)) (Fig. 5b). The band at 1661 cm\(^{-1}\) (Fig. 6f) is due to the C=C stretching vibration (\(\nu(C=C)\)). When CO is conjugated with C=C, \(\pi\) electron delocalization would occur, and weaken C=O double bond characteristic, then make the band shift to the low wavenumber. Hence, there may be methyl vinyl ketone (CH\(_3\)COCH = CH\(_2\)), isopropyl acrylate (CH\(_2\)=CHCOOCH\((CH_3)\)\(_2\)), and 2, 4-dimethylfuran (Fig. 5d). The band at 1446 cm\(^{-1}\) is attributed to symmetrical stretching vibration (\(\nu_s(COO)\)) of acetate species, indicating the formation of acetic acid (CH\(_3\)COOH) and at 1094 cm\(^{-1}\) is due to the OCC\(_{ee}\) stretching vibration (\(\nu(OCC_{ee})\)) of secondary alcohol, indicating the existence of isopropyl acetate (CH\(_3\)COOCH\((CH_3)\)\(_2\)). As for the O\(_e\)CC\(_e\) stretching vibration of alcohol, here we attribute the absorption band at 1094 cm\(^{-1}\) to the OCC\(_{ee}\) stretching vibration of secondary alcohol ester, because the O–CC– vibration band in ester is related to the type of ester. The O–CC– band of primary alcohol ester is 1064–1031 cm\(^{-1}\), while that of secondary alcohol ester is 1100 cm\(^{-1}\) [38]. The band at 1015 cm\(^{-1}\) is due to the C–O stretching vibration (\(\nu(C–O)\)) of ethers (R–O–R') (Fig. 6d and Fig. 6f). The band at 1193 cm\(^{-1}\) is due to the C(=O)‒O‒C asymmetric stretching vibration (\(\nu_{as}(C(=O)–O–C)\)) of propionate (CH\(_3\)CH\(_2\)COOH(CH\(_3\))\(_2\)). To be exact, the present propionate is isopropyl propionate (CH\(_3\)CH\(_2\)COOCH\((CH_3)\)\(_2\)) detected by GC–MS.

It should be noted that although high acetonel selectivity and yield could be achieved over 1.36 wt% Au/α-Fe\(_2\)O\(_3\) NSs under certain reaction conditions, the possible little residual products, including propylene, acetic acid, acetaldehyde, methyl vinyl ketone, 2-butanone, isopropyl ether, isopropyl acetate, 3-penten-2-one, isopropyl acrylate, isopropyl propionate, and 2, 4-dimethylfuran, would cause harmful effect on human health (Table S5). Therefore, it is important to take necessary measures (e.g. activated carbon adsorption) to remove these products.

3.5. Adsorption, activation, and transformation behaviors of isopropanol over commercial α-Fe\(_2\)O\(_3\), α-Fe\(_2\)O\(_3\) NSs, and x wt% Au/α-Fe\(_2\)O\(_3\) NSs

In order to investigate the adsorption, activation, and transformation behaviors of isopropanol, we carry out the isopropanol temperature programmed desorption (isopropanol-TPD) (Fig. 7) and isopropanol temperature programmed surface reaction (isopropanol-TPSR) (Fig. 8) experiments over α-Fe\(_2\)O\(_3\) NSs and Au/α-Fe\(_2\)O\(_3\) NSs. Before the isopropanol-TPD experiment, α-Fe\(_2\)O\(_3\) NSs and Au/α-Fe\(_2\)O\(_3\) NSs are pretreated by isopropanol pulse adsorption at room temperature, and then purged by 30 mL/min of He flow from room temperature to 900 °C. The fragment ions are detected by on-line mass spectrometry. Table 4 summarizes the relative strength of fragments for the possible compounds. Specially, the relative strength of fragments for isopropanol decreases in the order of \(m/z = 45 \gg m/z = 43 \gg m/z = 41 \gg m/z = 58\) (Fig. S8). For isopropanol ((CH\(_3\))\(_2\)CHOH) and acetone (CH\(_3\)COCH\(_3\)), the most abundant fragment is \(m/z = 45\) and \(m/z = 43\), respectively. Clearly, over α-Fe\(_2\)O\(_3\) NSs and Au/α-Fe\(_2\)O\(_3\) NSs, the desorption peak of (CH\(_3\))\(_2\)CHOH begins to appear at 66–237 °C. Another characteristic fragment for CH\(_3\)COCH\(_3\) is \(m/z = 58\). The desorption profiles of \(m/z = 43\) and \(m/z = 58\) exhibit similar pattern.

It should be noted that in addition to acetone, a number of intermediate
products including acetic acid (CH₃COOH), 2-butanone (CH₃COCH₂CH₃), isopropyl acetate (CH₃COOCH(CH₃)₂), methyl vinyl ketone (CH₃COCH = CH₂), 3-penten-2-one (CH₃COCH = CHCH₃), acetaldehyde (CH₃CHO), isopropyl propionate (CH₃CH₂COOCH(CH₃)₂), isopropyl acrylate (CH₂=CHCOOCH(CH₃)₂), and isopropyl ether ((CH₃)₂CHOHC(CH₃)₂) also could produce large quantity of fragment of m/z = 43 (Table 4). In other words, the desorption peak of m/z = 43 detected in the range of room temperature to 300 °C could not ignore the contribution of intermediate product desorption. The fragment of m/z = 41 might be the characteristic ion fragment of 3-penten-2-one (CH₃COCH = CHCH₃) and/or propylene (CH₂=CHCH₃). As shown in Fig. S9, the propylene selectivity over α-Fe₂O₃ NSs and 1.36 wt% Au/α-Fe₂O₃ NSs is less than 2%. At T90%, 3-penten-2-one is clearly observed over α-Fe₂O₃ NSs, and not found over 1.36 wt% Au/α-Fe₂O₃ NSs (Fig. 5b and Fig. 5d). Therefore, the weak desorption peak of m/z = 41 over Au/α-Fe₂O₃ NSs is mainly due to the propylene desorption, while that over α-Fe₂O₃ NSs is due to the 3-penten-2-one (CH₃COCH = CHCH₃) and/or propylene desorption. The fragment of m/z = 18 and m/z = 44 might be the characteristic ion fragment of H₂O and CO₂, respectively. Over α-Fe₂O₃ NSs and Au/α-Fe₂O₃ NSs, a
sharp desorption peak of m/z = 18 is detected at 68 °C, due to the selective catalytic oxidation of isopropanol, while that is detected in the range of 150–600 °C, due to the complete catalytic oxidation of reactants and products. In addition, the desorption peak of m/z = 44, due to the complete catalytic oxidation, is mainly observed in the range of 300–600 °C. Such results are consistent with the results of catalytic performance (Figs. S5, S6 and S9). Over the present catalysts and under the adopted reaction conditions, selective catalytic oxidation mainly occurs in the low temperature region, while complete catalytic oxidation mainly occurs in the high temperature region.

Before the isopropanol-TPSR experiment, commercial α-Fe2O3, α-Fe2O3 NSs and Au/α-Fe2O3 NSs are pretreated by isopropanol pulse adsorption at room temperature, and then purged by 30 mL/min of 20 % O2-80 % He (v/v) mixture from room temperature to 500 °C. Isopropanol-TPSR results are shown in Fig. 8 and Fig. S10. Over commercial α-Fe2O3, the strong peaks are detected at m/z = 45, 41, 44, and 18, as well as at m/z = 59, 57, and 55. The weak peaks are detected at m/z = 43 and 58. As shown in Fig. S7 and Fig. 5c, propylene (CH2=CHCH3) is the main product, in addition to little acetone and isopropyl ether ((CH3)2CHOHC(CH3)2) is generated from the intermolecular dehydration reaction of isopropanol ((CH3)2CHOH), while isopropyl ether ((CH3)2CHOHC(CH3)2) is generated from the intermolecular dehydration reaction of isopropanol. Under certain reaction conditions, propylene (CH2=CHCH3) could be partially oxidized into acrylic acid (CH2=CHCOOH) or propionic acid (CH3CH2COOH). Isopropyl acrylate (CH3CH2COOCOOH(CH3)3) and isopropyl propionate (CH3CH2COOOCOCH(CH3)3) might be generated from the esterification reaction between acrylic acid (CH2=CHCOOH) or propionic acid (CH3CH2COOH) and isopropanol ((CH3)2CHOH). 3-Penten-2-one (CH3COCH=CH2)/isopropyl acrylate (CH2=CHCOOCH(CH3)2) is the main product, in addition to little acetone and isopropyl ether ((CH3)2CHOHC(CH3)2) might be generated from the esterification reaction of isopropanol. Under certain reaction conditions, propylene (CH2=CHCH3) could be partially oxidized into acrylic acid (CH2=CHCOOH) or propionic acid (CH3CH2COOH). Isopropyl acrylate (CH3CH2COOCOOH(CH3)3) and isopropyl propionate (CH3CH2COOOCOCH(CH3)3) might be generated from the esterification reaction between acrylic acid (CH2=CHCOOH) or propionic acid (CH3CH2COOH) and isopropanol ((CH3)2CHOH). 3-Penten-2-one

### Table 4

The relative strength of fragment ions for different compounds.

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<tr>
<th>Compound</th>
<th>Molecular formula</th>
<th>m/z 18</th>
<th>m/z 41</th>
<th>m/z 43</th>
<th>m/z 44</th>
<th>m/z 45</th>
<th>m/z 55</th>
<th>m/z 57</th>
<th>m/z 58</th>
<th>m/z 59</th>
<th>m/z 69</th>
<th>m/z 72</th>
<th>m/z 87</th>
<th>m/z 96</th>
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<td>100</td>
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<td>3</td>
<td>100</td>
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<tr>
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<td>100</td>
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3.6. The possible reaction mechanism for the selective catalytic oxidation of isopropanol to acetone over commercial α-Fe2O3, α-Fe2O3 NSs, and x wt% Au/α-Fe2O3 NSs

Based on the results of GC–MS (Fig. 5), in situ DRIFTS (Fig. 6), isopropanol-TPSR (Figs. 7 and S8), and isopropanol-TPD (Figs. 8 and S10), we propose the possible reactions occur during the selective catalytic oxidation of isopropanol to acetone over the present catalysts (Scheme 1). Although α-Fe2O3 NSs and 1.36 wt% Au/α-Fe2O3 NSs exhibit high selectivity to acetone, a number of products are detected. Propylene (CH2=CHCH3) is generated from the intramolecular dehydration reaction of isopropanol ((CH3)2CHOH), while isopropyl ether ((CH3)2CHOH(CH3)2) is generated from the intermolecular dehydration reaction of isopropanol. Under certain reaction conditions, propylene (CH2=CHCH3) could be partially oxidized into acrylic acid (CH2=CHCOOH) and propionic acid (CH3CH2COOH). Isopropyl acrylate (CH3:CH=COOCOOH(CH3)3) and isopropyl propionate (CH3:CH2COOOCOCH(CH3)3) might be generated from the esterification reaction between acrylic acid (CH2=CHCOOH) or propionic acid (CH3CH2COOH) and isopropanol ((CH3)2CHOH). 3-Penten-2-one

[Scheme 1. Possible reactions occur during the selective catalytic oxidation of isopropanol to acetone over the present catalysts.]
(CH₃COCH = CHCH₃) is generated from the aldol condensation reaction between acetaldehyde (CH₃CHO) and acetone (CH₃COCH₃). Isopropyl acetate (CH₃COOCH(CH₃)₂) is generated from the esterification reaction between acetic acid (CH₃COOH) and isopropanol ((CH₃)₂COH). The isomerization of acetone (CH₃COCH₃) could generate less stable allyl alcohol (CH₂=COCH₃). On one hand, allyl alcohol might be further oxidized into small molecules, including acetaldehyde (CH₃CHO), acetic acid (CH₃COOH), and formaldehyde (HCHO). On the other hand, the keto-enol tautomerism of acetone (CH₃COCH₃) and allyl alcohol (CH₂=COCH₃) would cause the generation of 2, 4-dimethylfuran. Methyl vinyl ketone (CH₃COCH = CH₂) might be obtained from the condensation reaction between formaldehyde (HCHO) and acetone (CH₃COCH₃), and 2-butanone (CH₃COCH₂CH₃) might be obtained from the selective hydrogenation of methyl vinyl ketone (CH₃COCH = CH₂). Of course, the proposed reaction mechanism needs further investigation, considering the presence of various active species (e.g. CH₂ = C(CH₃) = O(ads), CH₃CHO*, CH₂O* and COO(ads)) generated from the catalytic oxidation process [41].

3.7. Influence of reactant composition and moisture on catalytic activity, selectivity, and stability of 1.36 wt% Au/α-Fe₂O₃ NSs

Considering the possible application background in the semiconductor industry, isopropanol is directly converted to acetone and then recycled. The 0.1 vol% concentration of isopropanol is relatively low, and far away from the practical demand. On the other hand, the explosive limit of isopropanol is in the range of 2–12 vol%. On the basis of the balance between demand and safety, we further evaluate the catalytic performance of 1.36 wt% Au/α-Fe₂O₃ NSs for selective catalytic oxidation of isopropanol under the reactants consisted of 0.75 vol % or 1.2 vol% isopropanol, and 40 vol% or 10 vol% O₂. Generally, more isopropanol and less O₂ would depress the complete oxidation reaction to some extent. 1.36 wt% Au/α-Fe₂O₃ NSs does not exhibit much difference in the catalytic performance in the presence of 0.75 vol% isopropanol and 40 vol% O₂, 1.2 vol% isopropanol and 40 vol% O₂, 0.75 vol% isopropanol and 10 vol% O₂, or 1.2 vol% isopropanol and 10 vol% O₂ (Fig. S9). In the range of 60–220 °C, isopropanol conversion increases with a rise in temperature, and the temperature required for 90 % (T90 %) conversion of isopropanol is about 180 °C. The major product is acetone, and acetone selectivity is more than 90% even at 220 °C. Under the condition of different reactant compositions, a small amount of propylene is detected in addition to acetone. The propylene selectivity over α-Fe₂O₃ NSs or 1.36 wt% Au/α-Fe₂O₃ NSs is less than 2% (Fig. S9C). In the high temperature reaction region, propylene selectivity over 1.36 wt% Au/α-Fe₂O₃ slightly increases with a rise in isopropanol concentration, and a decrease in O₂ concentration. In the presence of 1.2 vol% isopropanol and 40 vol% O₂, acetone selectivity and yield over 1.36 wt% Au/α-Fe₂O₃ NSs could be as high as about 99 % and 95 % at 220 °C, respectively.

We also investigate the moisture effect. Generally, isopropanol conversion, acetone selectivity and yield, as well as CO₂ selectivity at a reaction temperature higher than 90 °C do not exhibit much difference at different relative humidities of 25 %, 50 %, 75 %, and 100 % over 1.36 wt% Au/α-Fe₂O₃ NSs (Fig. 9). At a low reaction temperature of 60 °C (Table 5), the presence of little water vapor (relative humidity = 25 %, 50 %, and 75 %) enhances the conversion of isopropanol, while the presence of large water vapor (relative humidity = 100 %) obviously decreases the conversion of isopropanol due to the competitive adsorption. At 60 °C, TOF_Au under 0 – 75 % of relative humidity is in the range of 44.9 – 68.6 (× 10⁻³ s⁻¹), larger than that (7.1 × 10⁻³ s⁻¹) under 100 % of relative humidity. Isopropanol consumption rate and acetone formation rate also show similar changing trend to TOF_Au. It should be noted that at a reaction temperature lower than 250 °C, acetone selectivity could be kept more than 90 %, and CO₂ selectivity is less than 5%. However, acetone selectivity greatly decreases and CO₂ selectivity obviously increases at 280 °C.

Catalytic stability is another important factor for the practical application. The present selective catalytic oxidation of isopropanol is carried out in an oxygen rich atmosphere, and maintaining the acetone selectivity and yield at a high level is relatively difficult. We herein further investigate the catalytic stability of 1.36 wt% Au/α-Fe₂O₃ NSs under various reaction temperatures and reactant compositions. In the presence of 0.1 vol% isopropanol and 40 vol% O₂ (Fig. S11a), isopropanol conversion decreases from 60 % to 50 %, acetone selectivity slightly decreases from 96 % to 91 %, and acetone yield decreases from 58 % to 46 %, after continues reaction at 45 °C (the temperature required for 60 % conversion of isopropanol, T60 %) for 10 h. The possible reason for the decrease in conversion, selectivity, and yield is due to the competitive adsorption of isopropanol and acetone at low temperatures. Actually, the isopropanol conversion, acetone selectivity and yield well remain after continuous reaction at 120 °C (the temperature required for 90 % conversion of isopropanol, T90 %) for 10 h (Fig. S11b). From Fig. S11c and Fig. S11d, the catalytic stability of 1.36 wt% Au/α-Fe₂O₃ NSs at ca. T90 % is very good within 30 h on-stream reaction time, in the presence of 0.75 vol% isopropanol and 10 vol% O₂, or 1.2 vol% isopropanol and 40 vol% O₂. High acetone selectivity (95–97 %) and yield (93–95 %) could be obtained even at 220 °C under 100 % of relative humidity (Fig. 10). Usually, particle aggregation, carbon deposition, crystal phase transformation, and change in surface composition could cause the deactivation of a supported noble metal catalyst. Therefore, we believe that the crystal structure and surface composition of the 1.36 wt% Au/α-Fe₂O₃ NSs catalyst would not be changed under the adopted reaction conditions.

3.8. The possible reasons for the better catalytic performance of 1.36 wt% Au/α-Fe₂O₃ NSs

Isopropanol could be converted to propylene by dehydration, and to acetone by dehydrogenation or oxidative dehydrogenation. Previously, it was found that isopropanol dehydrogenation could happen over anatase TiO₂, and increasing surface –OH and oxygen partial pressure benefited dehydrogenation rate, while not affected dehydration rate [42]. The deposition of Pd onto the surface of TiO₂ might change the reaction pathway of propanol decomposition, and caused the generation of propylene and acetone [43]. Under the present reaction conditions, the dehydration reaction ((CH₃)₂CHOH → CH₂ = C(CH₃) = O(ads), CH₃CHO*) mainly proceeds over commercial Fe₂O₃ (Fig. S7), while the oxidative dehydrogenation reaction ((CH₃)₂COH + O₂ → CH₃COCH₃ + H₂O) mainly proceeds over α-Fe₂O₃ NSs and Au/α-Fe₂O₃ NSs (Figs. S5 and S9) [44]. It has been reported that strong acidic and weak basic sites benefit the dehydration of isopropanol to propylene, while moderate acidic and strong basic sites benefit the dehydrogenation of isopropanol to acetone [44,45]. Based on the results of NH₃–TPD and pyridine-in situ DRIFTS experiments (Figs. 3 and 4, Table 2), we find that the NH₃-desorption temperature decreases, and the NH₃ desorption amount increases in order of commercial α-Fe₂O₃, α-Fe₂O₃ NSs, and 1.36 wt% Au/α-Fe₂O₃ NSs, and the ratio of Bronsted acidic sites to Lewis acidic sites decreases in order of α-Fe₂O₃ NSs, commercial α-Fe₂O₃, and 1.36 wt% Au/α-Fe₂O₃ NSs. The NH₃ desorption temperature and amount, as well as the ratio of Bronsted acidic sites to Lewis acidic sites, indicate the acid strength and amount. In other words, commercial α-Fe₂O₃ possesses more strong acidic sites, while α-Fe₂O₃ NSs and 1.36 wt% Au/α-Fe₂O₃ NSs possess more moderate acidic sites. In addition to the acidic sites, the redox property might also play an important role in the selective or complete catalytic oxidation of propylene [45,46]. Liu and Yang [45] investigated the complete oxidation of isopropanol over CeO₂, Fe₂O₃, TiO₂, and Al₂O₃ supported Au catalysts, and found that Au/CeO₂ indicated higher activity due to its better redox and basic properties. The loading of Au NPs onto the surface of CeO₂ weakened the bonds of Ce – O nearby Au atoms, and then increased the mobility of the surface reactive lattice oxygen species. Gong et al. [46] reported
that pre-covered atomic oxygen on Au(111) could act as a Brønsted base for the first hydrogen abstraction. The dehydrogenation of isopropanol proceeded according to the cleavage of O–H bond, and then the selective activation of β-C–H bond. The cleavage of C–O bond did not happen, and the cleavage of non-selective γ-C–H bond was inhibited due to the atomic oxygen pre-covered Au(111). Finally, high selectivity to acetone could be achieved. The possible mechanism for molecular oxygen activation over the metal oxide supported Au catalyst involves O₂ dissociation on the surface oxygen vacancies, and the generation of oxygen atoms. We herein find that defect sites (including oxygen vacancy) increase on the support after the modification of Au NPs from the results of XRD (Fig. S1) and ⁵⁷Fe-Mössbauer (Fig. 1) characterizations. From the results of XPS characterization (Fig. S4 and Table 1), we find that the loading of Au NPs onto α-Fe₂O₃ NSs causes an interaction (Au° + Fe³⁺ → Auδ⁺ + Fe²⁺), and weakens the bonds of Fe–O nearby Au atoms. Furthermore, the surface O_ads/O_latt molar ratio increases from 0.24 to 0.38 with a rise in the Au loading from 0 to 1.36 wt %. Usually, the O_ads species are adsorbed on the oxygen vacancies [47]. In other words, the 1.36 wt% Au/α-Fe₂O₃ NSs catalyst possesses larger oxygen vacancies. The morphology of the support might influence the catalytic performance of Au/α-Fe₂O₃ nanocatalysts. The present α-Fe₂O₃ nanosheets support mainly expose the (110) crystal plane (Fig. S2), with larger Fe atom density and oxygen vacancy concentration [22,48–50]. From the results of H₂-TPR characterization (Fig. 2), the interaction between Au NPs and α-Fe₂O₃ NSs is further confirmed, since the temperature of the initial reduction peak shifts to low temperature with a rise in the Au loading. Therefore, the excellent catalytic performance of 1.36 wt% Au/α-Fe₂O₃ NSs for the selective catalytic oxidation of isopropanol to acetone is due to the presence of moderate acidic sites, larger surface oxygen vacancies, and better low-temperature reducibility [19,27,32,45,46].

4. Conclusions

Generally, we fabricate α-Fe₂O₃ NSs and their supported 0.38, 0.81, and 1.36 wt% Au nanocatalysts via the solvothermal and colloidal deposition method, respectively. The α-Fe₂O₃ nanocrystals display regular hexagonal morphology, and the width and thickness are about 180 and 15 nm, respectively. Au NPs (average particle size = 4.0 nm) are well dispersed on the surface of α-Fe₂O₃ NSs. Commercial α-Fe₂O₃ possesses the catalytic performance of Au/α-Fe₂O₃ nanocatalysts.

Table 5

<table>
<thead>
<tr>
<th>Relative humidity (%)</th>
<th>Isopropanol conversion (%)</th>
<th>Acetone selectivity (%)</th>
<th>Acetone yield (%)</th>
<th>( \text{TOF}_{\text{Au}} \times 10^{-3} \text{ s}^{-1} )</th>
<th>Isopropanol consumption rate ( \times 10^{-7} \text{ mol/(gcat s)} )</th>
<th>Acetone formation rate ( \times 10^{-7} \text{ mol/(gcat s)} )</th>
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</thead>
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<tr>
<td>0</td>
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<td>97.8</td>
<td>31.8</td>
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<td>8.6</td>
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<td>47.3</td>
<td>68.6</td>
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<td>93.3</td>
<td>37.5</td>
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<td>48.4</td>
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<tr>
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<td>92.3</td>
<td>4.8</td>
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Appendix A. Supplementary data

Supplementary material related to this article can be found, in the online version, at doi:https://doi.org/10.1016/j.apcatb.2020.119011.

References


5. Credit author statement
Hongzhong Zhang carried out experiments, wrote and modified the manuscript.
Lingyun Dai modified the manuscript.
Yuan Feng and Yuhan Xu carried out some experiments.
Yuxi Liu, Guangsheng Guo, Hongxing Dai, Chongchen Wang, Can Wang, Hsing-Cheng Hsi, and Haibao Huang analyzed the experimental results.
Jiguang Deng designed the experiments, analyzed the experimental results, wrote and modified the manuscript.

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